

Be sure to show all work, use the correct number of significant figures, circle final answers and use correct units in all problems.

1. Calculate these expressions. Include the correct number of significant figures and units in the answer. (4 points)

$$(9.994 \text{ g} - 8.33 \text{ g}) / (1.44 \text{ cm}^3 - 0.536 \text{ cm}^3) \quad \underline{\hspace{2cm}}$$

$$(9.16 * 10^{+3} \text{ mL}) * (2.3411 * 10^{-6} \text{ g}) / 12.001 * 10^{-3} \text{ g} \quad \underline{\hspace{2cm}}$$

2. Convert 892.0 °C to °F. (4 points)

3. Density Problem (8 points)

- a. Calculate the density of “substance A” when its mass = 11.22 g and its volume is 0.244 cm³. Express the density in units of g / mm³.
- b. It costs 10.79 cents to make 11.4 cm³ of “substance A”. Calculate how much it would cost to make 8.91 pounds of “substance A” in units of \$. (1 pound = 453.59237 g; 100 cents = 1 \$)

4. Differentiate between chemical and physical properties. Give at least one example of each. (4 points)